

I'm not robot!

1. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$	2. $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
3. $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$	4. $\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$
5. $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$	6. $\text{K}_3\text{PO}_4 + \text{Al}(\text{NO}_3)_3 \rightarrow 3\text{KNO}_3 + \text{AlPO}_4$

STOICHIOMETRY: MOLE-MOLE PROBLEMS

Name Abigail Stewart

1. $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
 How many moles of hydrogen are needed to completely react with two moles of nitrogen? $1:3:2 \times 2$
 $2:6:4$
6 moles of H_2

2. $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$
 How many moles of oxygen are produced by the decomposition of six moles of potassium chlorate?
 $2:3:2 \times 3$
 $6:9:6$
9 moles of O_2

3. $\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$
 How many moles of hydrogen are produced from the reaction of three moles of zinc with an excess of hydrochloric acid?
 $1:2 \rightarrow 1:1$
 $3:6 \rightarrow 3:3$
3 moles of H_2

4. $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$
 How many moles of oxygen are necessary to react completely with four moles of propane (C_3H_8)?
 $1:5 \rightarrow 3:4$
 $4:20 \rightarrow 12:16$
20 moles of O_2

5. $\text{K}_3\text{PO}_4 + \text{Al}(\text{NO}_3)_3 \rightarrow 3\text{KNO}_3 + \text{AlPO}_4$
 How many moles of potassium nitrate are produced when two moles of potassium phosphate react with two moles of aluminum nitrate?
 $1:1 \rightarrow 3:1$
 $2:2 \rightarrow 6:2$
6 moles of KNO_3

Worksheet: Mole/Mole Problems

Name _____

Answer each of the following questions using the equation provided. BE SURE TO BALANCE EACH EQUATION BEFORE YOU BEGIN ANY PROBLEMS. SHOW ALL WORK.

I. $\text{NO} + \text{O}_2 \rightarrow \text{NO}_2$

a. 8 moles of NO will react with _____ mole(s) of O_2 to produce _____ mole(s) of NO_2 .

b. 3 moles NO_2 = 3.6 moles O_2 = $\frac{\text{mole NO}_2}{\text{mole O}_2}$

c. How many moles of NO must react to form 4.07 moles of NO_2 ?

II. $\text{NH}_3 + \text{O}_2 \rightarrow \text{N}_2 + \text{H}_2\text{O}$

a. 20 moles of NH_3 are needed to produce _____ moles of H_2O .

b. How many moles of N_2 will be produced if 3.5 moles of O_2 react?

Repeating a sequence of characters (likely a keyboard test or a specific character set) across multiple lines.

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